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Christopher Orlando Sylvester Mawson

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 $\text{Pb(NO}_3)_2(\text{aq}) + 2\text{HCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{HNO}_3$

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web answer key to practice problems on net ionic equations molecular $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$ total ionic $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$

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web the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in

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web the net ionic equation for equation 1 is given as equation 3 $\text{S}_2(\text{aq}) + \text{Fe}_2(\text{aq}) + \text{GT}(\text{fcs}) \rightleftharpoons \text{S}(\text{eq}) + 3\text{H}_2$ note that the phase indicator aq is optional on ionic species but is

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web $\text{Cu}(\text{ClO}_4)_2$ is soluble all perchlorates are soluble while PbSO_4 is insoluble sulfates are generally soluble except for Pb^{2+} Ba^{2+} Hg_2^{2+} and Sr^{2+} the net ionic equation is

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web balanced chemical equation $2\text{NaOH}(\text{aq}) + \text{MgCl}_2(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{Mg}(\text{OH})_2(\text{s})$ full ionic equation $2\text{Na}(\text{aq}) + 2\text{OH}(\text{aq}) + \text{Mg}(\text{aq}) + 2\text{Cl}(\text{aq}) \rightarrow 2\text{Na}(\text{aq}) + 2\text{Cl}(\text{aq})$

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