

$$\frac{5}{2} n(\text{MnO}_4^-) = n(\text{H}_2\text{O}_2(\text{aq}))$$

$$= \frac{5}{2} \times \frac{28.55}{10^3} \times 0.02$$

$$n(\text{H}_2\text{O}_2) = 1.4275 \times 10^{-3} \text{ mol. in } 10 \text{ cm}^3 \text{ sample}$$

$$n(\text{H}_2\text{O}_2) = 1.4275 \times 10^{-3} \times \frac{250}{10} \text{ in } 250 \text{ cm}^3 \text{ ~~not~~ diluted,}$$

$$= 0.0356875 \dots \text{ AND in } 25 \text{ cm}^3 \text{ undiluted.}$$

↑ can you see why?

$$\therefore c(\text{H}_2\text{O}_2) = \frac{0.0356875 \dots}{\left(\frac{25}{10^3}\right)} = 1.4275 \text{ mol dm}^{-3}$$

$$= 1.4275 \times M_r(\text{H}_2\text{O}_2) \text{ g dm}^{-3}$$

$$= 1.4275 \times (2+32) \text{ g dm}^{-3}$$

$$= 48.5 \text{ g dm}^{-3}$$

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**CL Gary**



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