

## Graphing Exercise #2

### Ideal Gas Law vs. van der Waals (Real gas) Equation Chem 401, Dr. Mahoney

Recall the Ideal Gas Law;

$$PV = nRT$$

For this equation to be relevant and accurate, we made a few assumptions about the behavior of the gas. We called this our Kinetic Molecular Theory and it stated:

1. The size of a particle is negligibly small
2. The average kinetic energy of a particle is proportional to the temperature
3. The particles engage in elastic collisions

Gases that don't behave ideally are called Real Gases. They don't follow the assumptions that we made in the Kinetic Molecular Theory. The Van der Waals Equation (or Real gas equation) is:

$$\left[ P + a \left( \frac{n}{V} \right)^2 \right] [V - nb] = nRT$$

The two terms that we introduced are  $a$  (with units of  $L^2 \cdot \text{atm} / \text{mol}^2$ ) and  $b$  (with units of  $L / \text{mol}$ ).

' $a$ ' corrects for the intermolecular forces that occur between molecules. If a molecule has strong intermolecular forces, which Kinetic Molecular Theory postulate does it violate?

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' $b$ ' corrects for the size of the molecule. Since not all molecules are extremely small, this will affect the volume term in the gas equations.

# Ideal Gas Graphing Exercise Answers

**M Mosston**



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