

## THE IDEAL GAS LAW

**T** = Kelvin temperature

See  
scratch  
paper

- How many moles of oxygen will occupy a volume of 2.50 liters at 1.20 atm and  $25^{\circ}\text{C}$ ?  
 $PV = nRT$   
 $n = \frac{PV}{RT} = \frac{1.20 \text{ atm} \cdot 2.50 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}} = 0.0122 \text{ mol}$
- What volume will 2.00 moles of nitrogen occupy at  $720 \text{ mmHg}$  and  $20^{\circ}\text{C}$ ?  
 $V = \frac{nRT}{P} = \frac{2.00 \text{ mol} \cdot 0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 293 \text{ K}}{0.947 \text{ atm}} = 5.00 \text{ L}$
- What pressure will be exerted by  $25.0 \text{ g}$  of  $\text{CO}_2$  at temperature of  $25^{\circ}\text{C}$  and a volume of  $500 \text{ mL}$ ?  
 $n = \frac{25.0 \text{ g}}{44.0 \text{ g/mol}} = 0.568 \text{ mol}$   
 $P = \frac{nRT}{V} = \frac{0.568 \text{ mol} \cdot 0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}}{0.500 \text{ L}} = 2.75 \text{ atm}$
- At what temperature will  $5.00 \text{ g}$  of  $\text{Cl}_2$  exert a pressure of  $900 \text{ mmHg}$  at a volume of  $750 \text{ mL}$ ?  
 $n = \frac{5.00 \text{ g}}{71.0 \text{ g/mol}} = 0.0704 \text{ mol}$   
 $T = \frac{PV}{nR} = \frac{1.18 \text{ atm} \cdot 0.750 \text{ L}}{0.0704 \text{ mol} \cdot 0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}} = 255 \text{ K}$
- What is the density of  $\text{NH}_3$  at  $800 \text{ mmHg}$  and  $25^{\circ}\text{C}$ ?  
 $V = 1.00 \text{ L}$   
 $n = \frac{PV}{RT} = \frac{1.05 \text{ atm} \cdot 1.00 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}} = 0.0434 \text{ mol}$   
 $\text{mass} = 0.0434 \text{ mol} \cdot 17.0 \text{ g/mol} = 0.738 \text{ g}$   
 $\text{density} = \frac{0.738 \text{ g}}{1.00 \text{ L}} = 0.738 \text{ g/L}$
- If the density of a gas is  $1.2 \text{ g/L}$  at  $745 \text{ mmHg}$  and  $20^{\circ}\text{C}$ , what is its molar mass?  
 $V = 1 \text{ L}$   
 $n = \frac{PV}{RT} = \frac{0.980 \text{ atm} \cdot 1.00 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 293 \text{ K}} = 0.0406 \text{ mol}$   
 $\text{mass} = 1.2 \text{ g/L} \cdot 1.00 \text{ L} = 1.2 \text{ g}$   
 $\text{molar mass} = \frac{1.2 \text{ g}}{0.0406 \text{ mol}} = 29.5 \text{ g/mol}$
- How many moles of nitrogen gas will occupy a volume of  $347 \text{ mL}$  at  $6680 \text{ mmHg}$  and  $27^{\circ}\text{C}$ ?  
 $n = \frac{PV}{RT} = \frac{8.79 \text{ atm} \cdot 0.347 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 300 \text{ K}} = 0.122 \text{ mol}$
- What volume will  $454 \text{ grams}$  (1 lb) of hydrogen occupy at  $1.05 \text{ atm}$  and  $25^{\circ}\text{C}$ ?  
 $n = \frac{454 \text{ g}}{2.02 \text{ g/mol}} = 225 \text{ mol}$   
 $V = \frac{nRT}{P} = \frac{225 \text{ mol} \cdot 0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}}{1.05 \text{ atm}} = 5140 \text{ L}$
- Find the number of grams of  $\text{CO}_2$  that exert a pressure of  $285 \text{ mmHg}$  at a volume of  $32.5 \text{ L}$  and a temperature of  $32^{\circ}\text{C}$ .  
 $n = \frac{PV}{RT} = \frac{1.03 \text{ atm} \cdot 32.5 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 305 \text{ K}} = 1.32 \text{ mol}$   
 $\text{mass} = 1.32 \text{ mol} \cdot 44.0 \text{ g/mol} = 58.1 \text{ g}$
- An elemental gas has a mass of  $10.3 \text{ g}$ . If the volume is  $58.4 \text{ L}$  and the pressure is  $758 \text{ mmHg}$  at a temperature of  $25^{\circ}\text{C}$ , what is the gas?  
 $n = \frac{PV}{RT} = \frac{1.00 \text{ atm} \cdot 58.4 \text{ L}}{0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K} \cdot 298 \text{ K}} = 2.38 \text{ mol}$   
 $\text{molar mass} = \frac{10.3 \text{ g}}{2.38 \text{ mol}} = 4.33 \text{ g/mol}$   
 $\text{gas} = \text{He}$

# Ideal Gas Law Answers Pg 24

**Lingjun Ying**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance *Gaseous Fuels*  
American Gas Association, 1948

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