Ideal Gas Law Problem #2

A sample of CO₂ with a mass of 0.250 g was placed in a 350. mL container at 127°C. What is the pressure exerted by the gas? (Solve for kPa).

$$P = X kPa$$

$$V = .350 L$$

$$R = 8.314 L/kPa$$

$$(.350 \text{ X}) = (.00568)(8.314)(400)$$

$$.350 X = 18.889048$$

$$X = 54.0 \text{ kPa}$$

$$T = 127^{\circ}C + 273$$

Ideal Gas Law Problems And Solutions Atm

Frederick R. Longo

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