Use the chart to determine the heat of formation for each reaction. Use formula  $\Delta$ \*Hinn= $\Sigma\Delta$ \*Hinducts -  $\Sigma\Delta$ \*Hinducts

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- 2H<sub>2</sub> + O<sub>2</sub> → 2H<sub>2</sub>O<sub>10</sub>
   d\*Hron= [2H2O] [2H2 + O2]
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- Δ\*Hrxn=--571.6 kJ

  2. CH<sub>4</sub> + 2O<sub>2</sub> → 2H<sub>2</sub>O<sub>10</sub> + CO<sub>2</sub>
  Δ\*Hrxn=[2H<sub>2</sub>O + CO<sub>2</sub>] − [CH<sub>4</sub> + 2O<sub>3</sub>]
  Δ\*Hrxn=[2(-285.8)+-393.5] − [-74.8+2(0)]
  Δ\*Hrxn=-890.3 kJ
- CuCl<sub>2</sub> + 2Ag<sub>O1</sub> → Cu<sub>D1</sub> + 2AgCl (ΔH\*<sub>1</sub> CuCl<sub>2</sub> = -215.6 kJ/mol)
   Δ\*Hrxn=[Cu<sub>D1</sub> + 2AgCl] [CuCl<sub>2</sub> + 2Ag<sub>O1</sub>]
   Δ\*Hrxn=[0 + 2[-127]] [-215.6 + 2[0]]
   Δ\*Hrxn= -38.4 kJ
- NaOH + HCl → H<sub>2</sub>O<sub>m</sub> + NaCl
   Δ\*Hrxn={H<sub>2</sub>O<sub>m</sub> + NaCl} [NaOH + HCl]
   Δ\*Hrxn=[-285.8 + -411.0] [-426.6 + -92.3]
   Δ\*Hrxn= -177.9 kJ
- 2NO + O<sub>2</sub> → 2NO<sub>2</sub>
   Δ\*Hrxn=[2NO<sub>2</sub>] [2NO + O<sub>2</sub>]
   Δ\*Hrxn=[2[33.9]] [2(90.4) + 0]
   Δ\*Hrxn=-113.0 kJ
- N<sub>2</sub>O + NO<sub>2</sub> → 3NO (ΔH\*, N<sub>2</sub>O = 82.6 kJ/mol)
   Δ\*Hron=[3NO] − [N<sub>2</sub>O + NO<sub>2</sub>]
   Δ\*Hron=[3(90.4)] − [82.6 + 33.9]
   Δ\*Hron= 154.7 kJ
- 7.  $CaC_2 + 2H_2O_{10} \rightarrow Ca(OH)_2 + C_2H_3$   $(\Delta H^*_1 CaC_2 = -62.8 \text{ kJ/mol})$   $\Delta^*Hron=[Ca(OH)_2 + C_2H_3] [CaC_2 + 2H_2O_{10}]$   $\Delta^*Hron=[-986.6 + 226.7] [-62.8 + 2(-285.8)]$   $\Delta^*Hron=-125.5 \text{ kJ}$

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but anything were to happen to hungary if me one summer would kill climbed a while visiting some near by mountain on the baths summit 67 she of 15 empress elizabeth had been built companion seeing a visitor

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