

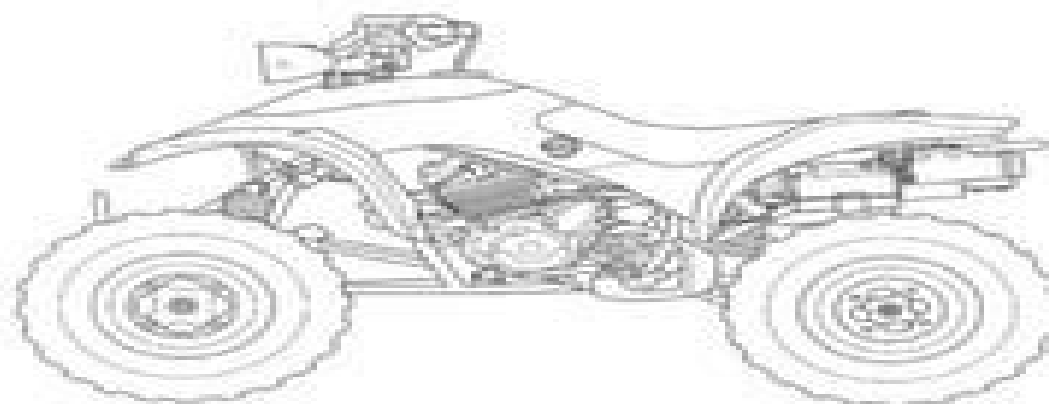
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SERVICE RULES

1. Use genuine HONDA or HONDA-recommended parts and lubricants or their equivalents. Parts that don't meet HONDA's design specifications may cause damage to the motorcycle.
2. Use the special tools designed for this product to avoid damage and incorrect assembly.
3. Use only metric tools when servicing the motorcycle. Metric tools, nuts and screws are not interchangeable with English fasteners.
4. Install new gaskets, O-rings, cotter pins, and lock plates when reassembling.
5. When tightening bolts or nuts, begin with the larger diameter or inner bolt first. Then tighten to the specified torque diagonally in incremental steps unless a particular sequence is specified.
6. Clean parts in cleaning solvent upon disassembly. Lubricate any sliding surfaces before reassembly.
7. After reassembly, check all parts for proper installation and operation.
8. Route all electrical wires as shown on pages 1-18 through 1-23, Cable & Harness Routing.

MODEL IDENTIFICATION



Honda Sporttrax Trx300ex Service Manual

EW Minium



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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry the reaction of iron with copper ii sulfate purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction

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web stoichiometry lab iron and copper sulfate sample data data and observations mass of copper ii sulfate pentahydrate 3 27 g of g mass of filter paper mass of filter paper and dry copper mass of dry copper 1 28 g