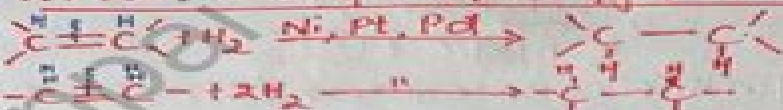


# HYDROCARBON

## # MOP of Alkane

### 1. From unsaturated hydrocarbon



DM  $\rightarrow$   $\pi$  bond ko H.H se tod do.

# exothermic

# 4 MCTS formed

# Syn addition

# Add<sup>n</sup> of conj. diene



#  $\text{RaNi} \rightarrow$  alloy of Ni & Al in NaOH  
Galvanom temp better than Ni (at 100°C)

# If  $\text{RaNi}$  used  $\rightarrow$  Sabatier sendraeson's

#  $[\text{Rh}(\text{PPh}_3)_3\text{Cl}]$  used also

(Wilkinson's catalyst)

Alkene	Addition	Product
Cis	Syn	Erythro
Trans	Syn	Threo
Cis	Anti	Threo
Trans	Anti	Erythro

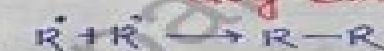
### 2. From alkyl halide



# X hatoo H lagao

#  $\text{ROR} : \text{R-Cl} < \text{R-Br} < \text{R-I}$   $\text{ROR} \propto \frac{1}{\text{BE}}$

### 3. Wurtz reaction



produce higher alkane with even no. of 'C'

#  $\text{R-X} \rightarrow$  coupling

#  $\text{R-X} \rightarrow$  disproportionation

# in 2 molecular wurtz faster than in 2 molecular but prod should be stable i.e A.A nahi banana hai.

#  $\text{ROR} \propto \frac{1}{\text{SH}}$   $\text{R-X} > 2^\circ \text{R-X} > 3^\circ \text{R-X}$

# for same R:  $\text{RF} < \text{RCI} < \text{RBr} < \text{RI}$

### 4. Decarboxylation reaction



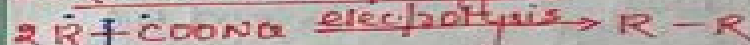
-COOH hatoo H lagao



-COONa hatoo H lagao

#  $\text{ROR} \propto$  stability of intermediate ( $\text{R}^\cdot$ )

### 5. Kolbe's electrolysis



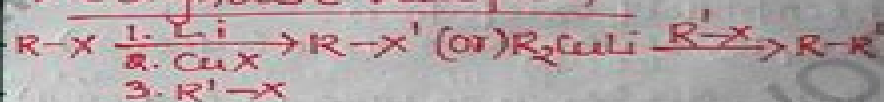
Cathode  $\rightarrow \text{H}_2$  Anode  $\rightarrow$  Alkane &  $\text{CO}_2$

### Addition Reaction

### 6. Frankland reaction



### 7. Coreyhouse reaction



# Attack of  $\text{R}^\cdot \rightarrow \text{RDS}$

#  $\text{ROR} \propto \frac{1}{\text{SH of R}'-\text{X}}$

# If same R:

$\text{R-F} < \text{R-Cl} < \text{R-Br} < \text{R-I}$

$\text{BE} \downarrow \rightarrow$  leaving tendency  $\uparrow$

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