

Instructor's Manual

**MATHEMATICAL
METHODS FOR
PHYSICISTS**
A Comprehensive Guide
SEVENTH EDITION

George B. Arfken
Miami University
Oxford, OH

Hans J. Weber
University of Virginia
Charlottesville, VA

Frank E. Harris
University of Utah, Salt Lake City, UT;
University of Florida, Gainesville, FL



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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} \rightarrow \text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ 2 make electrons equal

convert to basic solution $4\text{CN}^- + 2\text{Au} + \text{CN}^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} \rightarrow \text{O}_2 + 2\text{OH}^-$ 3 the final answer

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every atom of nickel oxidized but where do those electrons go

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