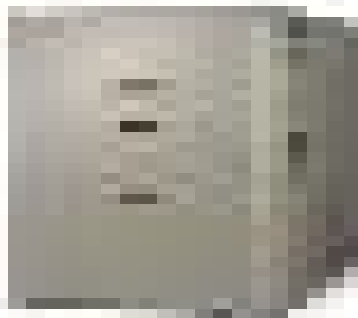


# LABORATORY

EXPERIMENT 1  
PURPOSE  
OBJECTIVES  
THEORY  
PROCEDURE



RESULTS  
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INTRODUCTION  
The purpose of this experiment is to determine the effect of temperature on the rate of reaction between hydrogen peroxide and potassium iodide. The reaction is exothermic and produces iodine and water. The rate of reaction is measured by the time taken for a fixed volume of iodine to be produced, which is indicated by the appearance of a blue color when starch is added as an indicator. The experiment is carried out at different temperatures and the results are plotted on a graph of log k against 1/T. The activation energy of the reaction is determined from the slope of the graph.

THEORY  
The rate of a chemical reaction is defined as the change in concentration of a reactant or product per unit time. The rate of reaction is affected by several factors, including temperature, concentration, and the presence of a catalyst. The Arrhenius equation relates the rate constant (k) of a reaction to the activation energy (Ea) and the absolute temperature (T):  
$$k = A e^{-E_a/RT}$$
  
where A is the pre-exponential factor, R is the gas constant, and T is the absolute temperature. Taking the logarithm of both sides of the equation gives:  
$$\log k = \log A - E_a/RT$$
  
A plot of log k against 1/T will give a straight line with a negative slope, from which the activation energy can be determined.

PROCEDURE  
1. Preparation of standard iodine solution: A known volume of a standard iodine solution is prepared by dissolving a precise amount of iodine in a known volume of potassium iodide solution. The concentration of the standard solution is determined by titration with a standard sodium thiosulfate solution.

2. Preparation of reaction mixture: A series of reaction mixtures are prepared by mixing different volumes of hydrogen peroxide, potassium iodide, and sulfuric acid. The total volume of the reaction mixture is kept constant.

3. Measurement of rate of reaction: The reaction is initiated by the addition of a small volume of the standard iodine solution to the reaction mixture. The time taken for a fixed volume of iodine to be produced is measured by the appearance of a blue color when starch is added as an indicator.

4. Plotting of graph: The rate of reaction (1/t) is plotted against the reciprocal of the absolute temperature (1/T). The activation energy is determined from the slope of the graph.

# Luxaire Furnace Manual

**Terry C. Jones**



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