

CHEM 141 Fall 2015

Lewis Structure Fundamentals

In previous lessons, we have investigated the **composition** of chemical compounds; that is, which elements are present in each, and in what atomic proportions. Knowledge of the chemical formula for a compound is important, but it is not the whole story in determining the identity of a compound (many may have the same chemical formula as one another!) or in predicting the properties of a compound. For a complete picture, we want to know not just the composition of a compound, but also its **constitution**; that is, which specific atoms are chemically bonded to one another in a molecule of the compound. There are many ways to represent the constitution of a compound; we will start in this course with the Lewis structure (named after Gilbert N. Lewis, who originated the idea of the covalent bond as a way for atoms to complete their valence electron shells without forming ions). The Lewis structure is the most detailed structural convention that chemists tend to use – the position of every atom and every valence electron (whether bonding or non-bonding) is accounted for explicitly, as are any non-zero formal charges present in the molecule. In this exercise, we will practice drawing reasonable Lewis structures for a variety of compounds.

Some general tips for drawing Lewis structures consisting mostly of p-block elements are as follows:

- Count the valence electrons in the compound based on the chemical formula and the number of valence electrons each neutral atom should contribute.
 - Adjust our valence electron count appropriately if your compound is charged. Add one electron for a negative charge, and subtract one for a positive charge.
- Draw a reasonable skeletal structure showing which atoms are connected to one another, using single bonds only.
 - Elements with lower group number or lower electronegativity tend to be central atoms.
 - Hydrogen atoms are virtually always terminal atoms, because they can only accommodate two electrons (a single bond).
 - Subtract the number of electrons in your skeletal structure (two for each bond) from the total number of valence electrons you have to work with to see how many remain.
- Add electrons to peripheral atoms to complete their octets.
 - Subtract the number of electrons added in this way from your electron tally.
 - If you still have valence electrons that need placing, add them to the central atom.
- Once all valence electrons have been accounted for, check that all atoms have complete valence shells (duets for hydrogen, octets for other p-block elements).
 - If any atom does not have a complete octet, you might need to use lone pairs on adjacent atoms to form double or triple bonds to complete the valence shell.

Give it a try! Go through the process with the compound phosphorus trifluoride (PF_3).

PF_3 : Total Valence Electrons: 26

Skeletal Structure	Complete Peripheral Octets	Final Structure
$\text{F}-\text{P}-\text{F}$ $\quad $ $\quad \text{F}$	$:\ddot{\text{F}}-\text{P}-\ddot{\text{F}}:$ $\quad $ $\quad :\ddot{\text{F}}:$	$:\ddot{\text{F}}-\ddot{\text{P}}-\ddot{\text{F}}:$ $\quad $ $\quad :\ddot{\text{F}}:$
Valence Electrons Remaining: 20	Valence Electrons Remaining: 2	Valence Electrons Remaining: ZERO

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