Equilibrium Involving Thiocyanatoiron III Ion

$$Fe(H_2O)_6^{3+} \text{ (Amber)} + SCN^{1-} \leftarrow \rightarrow Fe(SCN)(H_2O)_5^{2+} \text{ (D. Red)}$$

$$solution containing [Fe(SCN)(H_2O)_5]^{2+}$$

$$[Fe(SCN)(H_2O)_5]^{3+}$$

- a) Add KCI (Fe³⁺ forms an insoluble ppt [FeCl₃] with Cl¹⁻)
 Result: Amber Colour
- b) Add Fe(NO₃)₃ Possible Stresses: Fe⁺³ and NO₃ Result: Darker Red
- c) Add KSCN Possible Stresses: K⁺ and SCN⁻ Result: Darker Red
- d) Add NaOH (Fe³⁺ forms ins. ppt [Fe(OH)₃] with OH¹⁻)
 Result: First Orange and then Amber Colour

Kc Of Iron Thiocyanate

Aditya Jangid, Arshdeep Kaur

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