

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PerV=nRT$ ", and the universal gas constant $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 kPa / 1 atm$ to get
 $R = 8.31 kPa \cdot L / (K \cdot mole)$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 atm)(12 L)}{4 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 K$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of $87^{\circ}C$, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 atm \cdot 31 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}$$

$$n = 1.2586 mol$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 K}{60 L}$$

$$P = 1.642 atm$$

$$or \quad P = 166.29 kPa$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of $56^{\circ}C$, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 327 K}{0.09 atm}$$

$$V = 2310.93 L$$

- 5) If I have 17 moles of gas at a temperature of $67^{\circ}C$, and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 390 K}{88.89 L}$$

$$P = 5.34 atm$$

$$or \quad P = 540.61 kPa$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 atm \cdot 25 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 K}$$

$$n = 0.5075 mol$$

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Table of Contents Ideal Gas Law Packet Answer Key

1. Understanding the eBook Ideal Gas Law Packet Answer Key
 - The Rise of Digital Reading Ideal Gas Law Packet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ideal Gas Law Packet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ideal Gas Law Packet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ideal Gas Law Packet Answer Key
 - Personalized Recommendations
 - Ideal Gas Law Packet Answer Key User Reviews and Ratings
 - Ideal Gas Law Packet Answer Key and Bestseller Lists
5. Accessing Ideal Gas Law Packet Answer Key Free and Paid eBooks

- Ideal Gas Law Packet Answer Key Public Domain eBooks
- Ideal Gas Law Packet Answer Key eBook Subscription Services
- Ideal Gas Law Packet Answer Key Budget-Friendly Options
- 6. Navigating Ideal Gas Law Packet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Ideal Gas Law Packet Answer Key Compatibility with Devices
 - Ideal Gas Law Packet Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ideal Gas Law Packet Answer Key
 - Highlighting and Note-Taking Ideal Gas Law Packet Answer Key
 - Interactive Elements Ideal Gas Law Packet Answer Key
- 8. Staying Engaged with Ideal Gas Law Packet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ideal Gas Law Packet Answer Key
- 9. Balancing eBooks and Physical Books Ideal Gas Law Packet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ideal Gas Law Packet Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ideal Gas Law Packet Answer Key
 - Setting Reading Goals Ideal Gas Law Packet Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ideal Gas Law Packet Answer Key
 - Fact-Checking eBook Content of Ideal Gas Law Packet Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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