

## Chapter 2

### Atoms, Molecules, and Ions

#### INSTRUCTOR'S NOTES

Although much of this chapter will be review for many students who have taken high school chemistry, the ideas included are so central to later study that class coverage will probably be necessary. Key topics are the structure of the atom and related information (atomic number, isotopes), the mole unit, the periodic table, chemical formulas and names, and the relationships between formulas and composition. Three to five class periods will probably be necessary in order to address the essentials in this chapter unless your students are well-versed in some of these topics.

Some points on which students have some problems or questions are:

- The rule of determining the charges on transition metal cations tells students that they can assume such ions usually have 2+ or 3+ charges (with 2+ charges especially prominent). They are often uneasy about being given this choice. We certainly emphasize that they will see other possibilities (and that even negative charges are possible but that they will not see them in the general chemistry course).
- Students have to be convinced that they have no choice but to learn the language of chemistry by memorizing the names and charges of polyatomic ions. They can be reminded that correct names and formulas are required to prevent serious consequences, such as the use of the wrong medicine which can have tragic results or the purchase of the wrong substance which leads to wasted resources.
- A very common problem students have is recognizing that  $\text{MgBr}_2$ , for example, is composed of  $\text{Mg}^{2+}$  and two  $\text{Br}^-$  ions. We have seen such combinations as  $\text{Mg}^{2+}$  and  $\text{Br}_2^{2-}$ .

#### SUGGESTED DEMONSTRATIONS

- Properties of Elements
  - Take as many samples of elements as possible to your lecture on the elements and the periodic table.
  - See the series by Alton Banks in the *Journal of Chemical Education* titled "What's the Use?" This series describes a different element each month and gives references to the *Periodic Table Videodisc*.
  - Pinto, G. "Using Balls from Different Sports to Model the Variation of Atomic Sizes," *Journal of Chemical Education* **1998**, 75, 725.
- Atomic Structure
  - Hohman, J. R. "Introduction of the Scientific Method and Atomic Theory to Liberal Arts Chemistry Students," *Journal of Chemical Education* **1998**, 75, 1578.
- Elements That Form Molecules in Their Natural States
  - Use samples of  $\text{H}_2$ ,  $\text{O}_2$ ,  $\text{N}_2$ , and  $\text{Br}_2$  to illustrate elements that are molecules.

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